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# WEEKLY PLANNER: ALL SCIENCE WEEK

## OF 5.6.19



Objectives for the week: REVIEW FOR EOC!!!

Day	Honors Biology	Honors Chemistry
Mon 5.6	<b>Kahoot! 3.1</b> <a href="https://www.youtube.com/watch?v=enVK-qWyDU0">https://www.youtube.com/watch?v=enVK-qWyDU0</a> <b>Species diversity</b>	-Potential Energy diagrams and $\Delta H$ -rxn rates <a href="https://www.youtube.com/watch?v=OttRV5ykP7A&amp;t=1s">https://www.youtube.com/watch?v=OttRV5ykP7A&amp;t=1s</a> -Radioactivity -Hand out sheet -Go over review 1-8 <a href="https://kahoot.it/challenge/0346980">https://kahoot.it/challenge/0346980</a>
Tues 5.7 <b>STUDY BUDDIES!</b>	<b>Kahoot! 3.1</b> <a href="https://kahoot.it/challenge/0456807">https://kahoot.it/challenge/0456807</a> <b>-students receive the 3.2 review</b>	<b>KAHOOT challenge!</b> <b>-REVIEW #9- present if needed!</b> -Class works on Released EOC <a href="https://kahoot.it/challenge/0798337">https://kahoot.it/challenge/0798337</a>
Wed 5.8 <b>STUDY BUDDIES!</b>	<b>Kahoot! 3.1</b> <a href="https://kahoot.it/challenge/0212127">https://kahoot.it/challenge/0212127</a> <b>Finish the 3.2 EOC review</b>	<b>KAHOOT challenge!</b> <b>Finish released EOC</b> <a href="https://kahoot.it/challenge/0213558">https://kahoot.it/challenge/0213558</a>

Thurs 5.9	Fair Bluff awards Ceremony -review	Fair Bluff awards ceremony -review
Friday 5.10	<b>LAST TEST</b> <b>3.3, 3.4, 3.5, 4.1,</b> <b>4.2</b> <b>DUE MON:</b> <a href="https://kahoot.it/challenge/0320062">https://kahoot.it/ challenge/03200 62</a>	<b>LAST TEST!</b> <b>OVER WEEKEND:</b> -5 tough gas law problems -3 tough Limiting reactant problems -2 empirical and molecular formula problems. -2 molarity and titration problems

**NO warm up this week... all beginning of the class activities will involve Kahoot reviews!**

#### OBJECTIVES AT A GLANCE- CHEMISTRY

Chm.1.1.1 Analyze the structure of atoms, isotopes, and ions. Chm.1.1.2 Analyze an atom in terms of the location of electrons. Chm.1.1.3 Explain the emission of electromagnetic radiation in spectral form in terms of the Bohr model. Chm.1.1.4 Explain the process of radioactive decay by the use of nuclear equations and half-life.

Chm.1.2.1 Compare (qualitatively) the relative strengths of ionic, covalent, and metallic bonds. Chm.1.2.2 Infer the type of bond and chemical formula formed between atoms. Chm.1.2.3 Compare inter- and intra- particle forces. Chm.1.2.4 Interpret the name and formula of compounds using IUPAC convention. Chm.1.2.5 Compare the properties of ionic, covalent, metallic, and network compounds.

Chm.1.3.1 Classify the components of a periodic table (period, group, metal, metalloid, nonmetal, transition). Chm.1.3.2 Infer the physical properties (atomic radius, metallic and nonmetallic characteristics) of an element based on its position on the Periodic Table. Chm.1.3.3 Infer the atomic size, reactivity, electronegativity, and ionization energy

Chm.2.1.1 Explain the energetic nature of phase changes. Chm.2.1.2 Explain heating and cooling curves (heat of fusion, heat of vaporization, heat, melting point, and boiling point). Chm.2.1.3 Interpret the data presented in phase diagrams. Chm.2.1.4 Infer simple calorimetric calculations based on the concepts of heat lost equals heat gained and specific heat. Chm.2.1.5 Explain the relationships between pressure, temperature, volume, and quantity of gas both qualitative and quantitative.

Chm.2.2.1 Explain the energy content of a chemical reaction. Chm.2.2.2 Analyze the evidence of chemical change. Chm.2.2.3 Analyze the law of conservation of matter and how it applies to various types of chemical equations (synthesis, decomposition, single replacement, double replacement, and combustion). Chm.2.2.4 Analyze the stoichiometric relationships inherent in a chemical reaction. Chm.2.2.5 Analyze quantitatively the composition of a substance (empirical formula, molecular formula, percent composition, and hydrates).

Chm.3.1.1 Explain the factors that affect the rate of a reaction (temperature, concentration, particle size and presence of a catalyst). Chm.3.1.2 Explain the conditions of a system at equilibrium. Chm.3.1.3 Infer the shift in equilibrium when a stress is applied to a chemical system (Le Chatelier's Principle).

Chm.3.2.1 Classify substances using the hydronium and hydroxide ion concentrations. Chm.3.2.2 Summarize the properties of acids and bases. Chm.3.2.3 Infer the quantitative nature of a solution (molarity, dilution, and titration with a 1:1 molar ratio). Chm.3.2.4 Summarize the properties of solutions. Chm.3.2.5 Interpret solubility diagrams. Chm.3.2.6 Explain the solution process.

## OBJECTIVES AT A GLANCE- BIOLOGY

Bio.1.1.1 Summarize the structure and function of organelles in eukaryotic cells (including the nucleus, plasma membrane, cell wall, mitochondria, vacuoles, chloroplasts, and ribosomes) and ways that these organelles interact with each other to perform the function of the cell. Bio.1.1.2 Compare prokaryotic and eukaryotic cells in terms of their general structures (plasma membrane and genetic material) and degree of complexity. Bio.1.1.3 Explain how instructions in DNA lead to cell differentiation and result in cells specialized to perform specific functions in multicellular organisms.

Bio.1.2.1 Explain how homeostasis is maintained in the cell and within an organism in various environments (including temperature and pH). Bio.1.2.2 Analyze how cells grow and reproduce in terms of interphase, mitosis and cytokinesis. Bio.1.2.3 Explain how specific cell adaptations help cells survive in particular environments (focus on unicellular organisms).

Bio.2.1.1 Analyze the flow of energy and cycling of matter (water, carbon, nitrogen and oxygen) through ecosystems relating the significance of each to maintaining the health and sustainability of an ecosystem. Bio.2.1.2 Analyze the survival and reproductive success of organisms in terms of behavioral, structural, and reproductive adaptations. Bio.2.1.3 Explain various ways organisms interact with each other (including predation, competition, parasitism, mutualism) and with their environments resulting in stability within ecosystems. Bio.2.1.4 Explain why ecosystems can be relatively stable over hundreds or thousands of years, even though populations may fluctuate (emphasizing availability of food, availability of shelter, number of predators and disease).

Bio.2.2.1 Infer how human activities (including population growth, pollution, global warming, burning of fossil fuels, habitat destruction and introduction of nonnative species) may impact the environment. Bio.2.2.2 Explain how the use, protection and conservation of natural resources by humans impact the environment from one generation to the next.

Bio.3.1.1 Explain the double-stranded, complementary nature of DNA as related to its function in the cell. Bio.3.1.2 Explain how DNA and RNA code for proteins and determine traits. Bio.3.1.3 Explain how mutations in DNA that result from interactions with the environment (i.e. radiation and chemicals) or new combinations in existing genes lead to changes in function and phenotype.

Bio.3.2.1 Explain the role of meiosis in sexual reproduction and genetic variation. Bio.3.2.2 Predict offspring ratios based on a variety of inheritance patterns (including dominance, co-dominance, incomplete dominance, multiple alleles, and sex-linked traits). Bio.3.2.3 Explain how the environment can influence the expression of genetic traits.

Bio.3.3.1 Interpret how DNA is used for comparison and identification of organisms. Bio.3.3.2 Summarize how transgenic organisms are engineered to benefit society. Bio.3.3.3 Evaluate some of the ethical issues surrounding the use of DNA technology (including cloning, genetically modified organisms, stem cell research, and Human Genome Project).

Bio.3.4.1 Explain how fossil, biochemical, and anatomical evidence support the theory of evolution. Bio.3.4.2 Explain how natural selection influences the changes in species over time. Bio.3.4.3 Explain how various disease agents (bacteria, viruses, chemicals) can influence natural selection.

Bio.3.5.1 Explain the historical development and changing nature of classification systems. Bio.3.5.2 Analyze the classification of organisms according to their evolutionary relationships (including dichotomous keys and phylogenetic trees).

Bio.4.1.1 Compare the structures and functions of the major biological molecules (carbohydrates, proteins, lipids, and nucleic acids) as related to the survival of living organisms. Bio.4.1.2 Summarize the relationship among DNA, proteins and amino acids in carrying out the work of cells and how this is similar in all organisms. Bio.4.1.3 Explain how enzymes act as catalysts for biological reactions.

Bio.4.2.1 Analyze photosynthesis and cellular respiration in terms of how energy is stored, released, and transferred within and between these systems. Bio.4.2.2 Explain ways that organisms use released energy for maintaining homeostasis (active transport).

## **BIO**

# **EOC ANSWERS and objectives to STUDY this**

# **weekend!**

1	MC	A	Bio.1.1.1
2	MC	A	Bio.1.1.3
3	MC	C	Bio.1.2.1
4	MC	B	Bio.1.2.2
5	MC	C	Bio.2.1.1
6	MC	A	Bio.2.1.3
7	MC	A	Bio.2.2.1
8	MC	A	Bio.2.2.2
9	MC	B	Bio.3.1.1
10	MC	D	Bio.3.1.2
11	MC	C	Bio.3.2.1
12	MC	B	Bio.3.2.1
13	MC	C	Bio.3.3.1
14	MC	D	Bio.3.3.2
15	TE	Online Only	Bio.3.4.1
16	MC	B	Bio.3.4.3
17	MC	C	Bio.3.5.1
18	MC	D	Bio.4.1.1
19	MC	A	Bio.4.1.2
20	MC	B	Bio.4.2.1
21	MC	D	Bio.1.1.1
22	MC	B	Bio.1.1.3
23	MC	C	Bio.1.2.1

24	MC	D	Bio.1.2.2
25	MC	D	Bio.2.1.2
26	MC	B	Bio.2.1.3
27	MC	A	Bio.2.2.1
28	MC	B	Bio.2.2.1
29	MC	C	Bio.3.1.2
30	MC	C	Bio.3.1.2
31	TE	Online Only	Bio.3.2.1
32	MC	B	Bio.3.2.2
33	MC	D	Bio.3.3.1
34	MC	C	Bio.3.3.2
35	MC	B	Bio.3.1.3
36	MC	D	Bio.3.4.3
37	MC	D	Bio.3.5.2
38	MC	C	Bio.4.1.1
39	MC	A	Bio.4.1.3
40	MC	A	Bio.4.2.2
41	MC	B	Bio.1.1.2
42	MC	D	Bio.1.1.3
43	MC	A	Bio.1.2.2
44	MC	C	Bio.1.2.3
45	MC	C	Bio.2.1.2
46	MC	D	Bio.2.1.3
47	MC	D	Bio.2.2.1
48	MC	B	Bio.2.2.2
49	TE	Online Only	Bio.3.1.1
50	MC	B	Bio.3.1.3
51	MC	B	Bio.3.2.1
52	MC	D	Bio.3.2.1
53	MC	B	Bio.3.3.1
54	MC	B	Bio.3.1.3
55	MC	A	Bio.3.4.2
56	MC	C	Bio.3.5.1
57	MC	C	Bio.3.5.2
58	MC	A	Bio.4.1.2
59	MC	A	Bio.4.2.1
60	MC	B	Bio.4.2.2

**CHEMISTRY NCFE released 2018**

1	MC	D	63%	Chm.1.1.1
2	MC	B	58%	Chm.1.1.2
3	MC	D	65%	Chm.1.1.3
4	MC	A	64%	Chm.1.1.4
5	MC	A	71%	Chm.1.1.1
6	MC	B	59%	Chm.1.1.2
7	MC	D	64%	Chm.1.2.1
8	MC	A	65%	Chm.1.2.2
9	MC	A	34%	Chm.1.2.3
10	MC	C	66%	Chm.1.2.4
11	MC	D	66%	Chm.1.2.5
12	MC	C	72%	Chm.1.2.4
13	MC	A	61%	Chm.1.2.5
14	MC	C	84%	Chm.1.3.1
15	MC	D	36%	Chm.1.3.2
16	MC	C	73%	Chm.1.3.3

17	MC	A	65%	Chm.1.3.2
18	MC	A	62%	Chm.2.1.1
19	MC	A	72%	Chm.2.1.3
20	MC	C	59%	Chm.2.1.4
21	MC	B	90%	Chm.2.1.5
22	MC	B	70%	Chm.2.1.3
23	MC	D	55%	Chm.2.1.5
24	MC	B	71%	Chm.2.2.1
25	MC	C	38%	Chm.2.2.2
26	MC	B	79%	Chm.2.2.3
27	MC	D	40%	Chm.2.2.4
28	MC	B	76%	Chm.2.2.5
29	MC	B	59%	Chm.2.2.2
30	MC	D	56%	Chm.2.2.4
31	MC	D	58%	Chm.3.1.1
32	MC	A	39%	Chm.3.1.3
33	MC	A	47%	Chm.3.1.2
34	MC	B	58%	Chm.3.1.3

35	MC	C	60%	Chm.3.2.1
36	MC	C	71%	Chm.3.2.2
37	MC	B	56%	Chm.3.2.3
38	MC	A	60%	Chm.3.2.4
39	MC	C	50%	Chm.3.2.5
40	MC	B	47%	Chm.3.2.6

43	B	2 — Structures & Properties of Matter
44	D	2 — Structures & Properties of Matter
45	A	2 — Structures & Properties of Matter
46	C	2 — Structures & Properties of Matter
47	D	2 — Structures & Properties of Matter
48	B	3 — Regularities in Chemistry
49	C	3 — Regularities in Chemistry
50	C	3 — Regularities in Chemistry
51	D	3 — Regularities in Chemistry
52	A	3 — Regularities in Chemistry
53	B	4 — Energy Changes in Chemistry
54	B	4 — Energy Changes in Chemistry
55	B	5 — Chemical Reactions
56	A	5 — Chemical Reactions
57	B	5 — Chemical Reactions
58	D	5 — Chemical Reactions
59	A	5 — Chemical Reactions
60	D	5 — Chemical Reactions



14	D	4 — Energy Changes in Chemistry
15	B	4 — Energy Changes in Chemistry
16	A	5 — Chemical Reactions
17	B	5 — Chemical Reactions
18	B	5 — Chemical Reactions
19	C	5 — Chemical Reactions
20	D	5 — Chemical Reactions
21	D	2 — Structures & Properties of Matter
22	B	2 — Structures & Properties of Matter
23	B	2 — Structures & Properties of Matter
24	A	2 — Structures & Properties of Matter
25	B	2 — Structures & Properties of Matter
26	C	2 — Structures & Properties of Matter
27	D	2 — Structures & Properties of Matter
28	A	3 — Regularities in Chemistry
29	C	3 — Regularities in Chemistry
30	A	3 — Regularities in Chemistry
31	C	3 — Regularities in Chemistry
32	C	3 — Regularities in Chemistry
33	C	4 — Energy Changes in Chemistry
34	D	4 — Energy Changes in Chemistry
35	B	4 — Energy Changes in Chemistry
36	C	5 — Chemical Reactions
37	D	5 — Chemical Reactions
38	B	5 — Chemical Reactions
39	A	5 — Chemical Reactions
40	A	5 — Chemical Reactions
41	C	2 — Structures & Properties of Matter
42	C	2 — Structures & Properties of Matter

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1 B  
2 D  
3 A  
4 C  
5 C  
6 C  
7 C  
8 D  
9 D  
10 A  
11 C  
12 B  
13 B  
14 C  
15 D  
16 B  
17 D  
18 A  
19 D  
20 D  
21 D  
22 C  
23 C  
24 A  
25 D  
26 A  
27 D  
28 D  
29 D  
30 A  
31 C  
32 A

33 B  
34 C  
35 C  
36 C  
37 D  
38 C  
39 D  
40 C  
41 B  
42 A  
43 C  
44 D  
45 A  
46 D  
47 B  
48 B  
49 C  
50 NA  
51 B  
52 C  
53 B  
54 D  
55 D  
56 C  
57 A  
58 C  
59 A  
60 B  
61 B  
62 C  
63 D  
64 B  
65 A  
66 D  
67 D  
68 B  
69 D  
70 C  
71 D  
72 C

73 A  
74 B  
75 C  
76 B  
77 C  
78 B  
79 D  
80 C  
81 D  
82 B  
83 D  
84 C  
85 A  
86 D  
87 D  
88 C  
89 C  
90 D  
91 B  
92 D  
93 A  
94 D