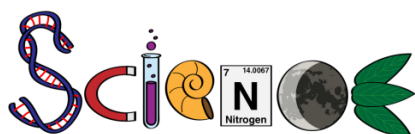


SCIENCE PLANNER: WEEK OF 11.4.19



OBJECTIVES FOR THE WEEK:

Biology : Bio.3.4.1 Explain how fossil, biochemical, and anatomical evidence support the theory of evolution. Bio.3.4.2 Explain how natural selection influences the changes in species over time. Bio.3.4.3 Explain how various disease agents (bacteria, viruses, chemicals) can influence natural selection.

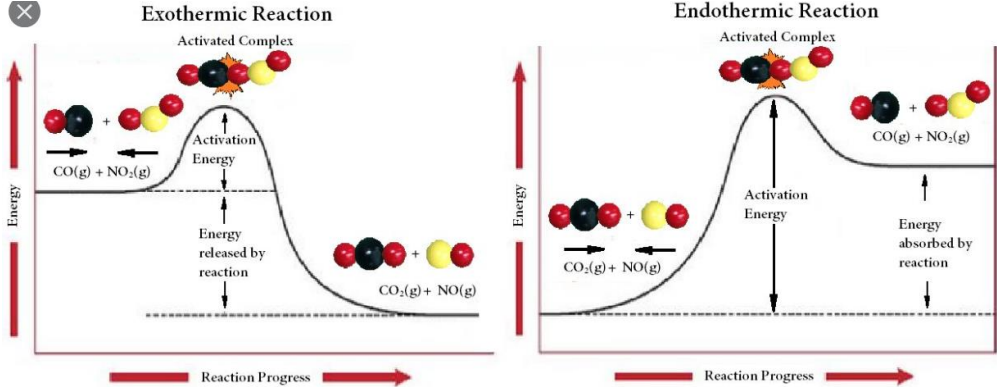
Chemistry: Chm.3.1.1 Explain the factors that affect the rate of a reaction (temperature, concentration, particle size and presence of a catalyst). Chm.3.1.2 Explain the conditions of a system at equilibrium. Chm.3.1.3 Infer the shift in equilibrium when a stress is applied to a chemical system (Le Chatelier's Principle).

DAILY AGENDA – (SUBJECT TO CHANGE) <https://evanscca.weebly.com/>

DAY	Honors Biology	Honors Chemistry
Mon 11.11		
Tues 11.12	<p>Notes: Change over Time / Evolution</p> <p>LAB: simple to complex</p> <p>HW= #1-16 in back of packet</p>	<p>EQUILIBRIUM mini LAB!</p> <p>https://www.youtube.com/watch?v=dUMmoPdwBy4</p> <p>Worksheet #'s ---</p> <p>https://www.youtube.com/watch?v=1GiZzCzm05Q</p> <p>HW= pg 79 and 80, finish paper wad lab, finish CK12!!</p>
Wed 11.13	<p>Change over time LAB!!</p> <p>https://www.pbs.org/wgbh/nova/labs/lab/evolution/</p> <p>Modules 1, 2 and 3</p> <p>*HW= TEST</p> <p>CORRECTIONS!!</p>	<p>https://www.youtube.com/watch?v=XmgRRmxS3is</p> <p>Be ready to turn in pg 79 and 80 to participate in the lab.</p> <p>QUIZ- Bobby, Bella and Steven</p> <p>LAB!!</p> <p>*HW= finish lab (#1 and 2 or 3 only) and test corrections</p>

<p>Thurs 11.7</p>	<p>Presentations of practice test!</p> <p>Finish lab modules</p> <p>HW= study for test.</p>	<p>Lechatelier notes: https://www.youtube.com/watch?v=XmgRRmxS3is Worksheets: Lechatellier principle. Pg 81 and 82.</p> <p>HW= Study for test, KAHOOT challenge: 0829020</p>
<p>Fri 11.8</p>	<p>TEST</p> <p>*HW= finish modules 1-3 on NOVA website 😊</p> <p>HAND IN MONDAY!!</p>	<p>TEST</p> <p>*HW= read chapter 6.3 and 6.4 on molarity and concentration:</p> <p>https://openstax.org/books/chemistry-atoms-first-2e/pages/6-3-molarity ANSWER the following 5 questions to hand in Monday:</p> <p>19. Explain what changes and what stays the same when 1.00 L of a solution of NaCl is diluted to 1.80 L.</p> <p>20. What information is needed to calculate the molarity of a sulfuric acid solution?</p> <p>21. A 200-mL sample and a 400-mL sample of a solution of salt have the same molarity. In what ways are the two samples identical? In what ways are these two samples different?</p> <p>22. Determine the molarity for each of the following solutions:</p> <p>(a) 0.444 mol of CoCl₂ in 0.654 L of solution</p> <p>(b) 98.0 g of phosphoric acid, H₃PO₄, in 1.00 L of solution</p>

WARM UP DISCUSSION QUESTIONS

MON	
TUES	<p>Describe the similarities and differences between a prokaryote and eukaryote.</p> <p>Describe the two reactions below:</p> <p>W</p>  <p>The image contains two energy profile diagrams. The left diagram is titled 'Exothermic Reaction' and shows a curve starting at a higher energy level for reactants (CO(g) + NO₂(g)) and ending at a lower energy level for products (CO₂(g) + NO(g)). The peak of the curve is labeled 'Activated Complex'. The energy difference between the reactants and products is labeled 'Energy released by reaction'. The right diagram is titled 'Endothermic Reaction' and shows a curve starting at a lower energy level for reactants (CO₂(g) + NO(g)) and ending at a higher energy level for products (CO(g) + NO₂(g)). The peak is labeled 'Activated Complex'. The energy difference is labeled 'Energy absorbed by reaction'. Both diagrams have 'Energy' on the vertical axis and 'Reaction Progress' on the horizontal axis.</p>
WED	<p>What is an example of change over time?</p> <p>Why do you NEVER use solids or liquids in an equilibrium constant expression??</p>
THU	<p>What are the evidences for evolution?</p> <p>What are examples of things that can shift dynamic equilibrium.</p>
FRI	<p>Draw and label two structures that are:</p> <ul style="list-style-type: none"> -homologous -analogous <p>0.02 M Ammonia solution is in dynamic equilibrium with .003M Nitrogen gas and</p>

	1.2×10^{-4} M Hydrogen gas. Write the balanced equation, the equilibrium constant expression, and solve for the constant. THEN explain if the products are favored.
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46. In the human kidney, urea from the blood is filtered through the glomerular membrane into a nephron. The movement of urea across this membrane occurs without an input of energy. Which factor is the MOST likely reason urea absorption does not require energy? A. a pH imbalance B. a pressure difference C. a temperature increase D. a concentration gradient

BIO A- 211

THUMB RULE!!

-answer with explanation

-picture

-anything else to help explain.