## SCIENCE PLANNER: WEEK OF10.21.19



## OBJECTIVES FOR THE WEEK:

Biology : What is DNA and how does it work? Bio.3.1.2 Explain how DNA and RNA code for proteins and determine traits. Bio.3.1.3 Explain how mutations in DNA that result from interactions with the environment (i.e. radiation and chemicals) or new combinations in existing genes lead to changes in function and phenotype.

Chemistry: How do we count (and account for) atoms? Chm.2.2.4 Analyze the stoichiometric relationships inherent in a chemical reaction.

DAILY AGENDA - (subject to change) https://evansccca.weebly.com/

| DAY | Biology | Chemistry |
| :---: | :---: | :---: |
| $\begin{aligned} & \hline \text { Mon } \\ & 10.21 \end{aligned}$ | NOTES: Heredity <br> Practice: Punnett Squares <br> *HW= MONOhybrid cross practice questions on your OWN paper. | THREE RING CIRCUS <br> -most students will be REDOING the test! <br> -The rest get to make the mole town maze!!! RY, BH, KJ, ..JB, AE, SLa <br> *HW= 16 balanced equations that match page 6 of reference table. |
| $\begin{array}{l\|} \hline \text { Tues } \\ 10.22 \end{array}$ | -hand in 10 cross practices -NOTES: dihybrid crosses *HW= quiz corrections and DO pg 7-10 of packet!! | -hand in 16 unique equations -TEST practice peeps -5 ALUMINUM equations (like warm up). -MOLETOWN ideas |
| $\begin{aligned} & \text { Wed } \\ & 10.23 \end{aligned}$ | Go over HW <br> Finish notes: complex genetics $\text { *HW= pg 21, } 25 \& 26$ | QUIZ <br> MOLE DAY CELEBRATION!!! <br> *HW=pg 1 of writing ionic compounds, do quiz corrections on Warm up!! |
| Thurs | WARM UP | QUIZ parts 3 \&4 |


| 10.24 | STATION LAB! <br> HW= double check ALL <br> homework, finish lab and <br> study for test. | Warm up will be quiz <br> corrections. <br> STATION LAB! <br> HW= finish lab and study for <br> test. |
| :--- | :--- | :--- |
| Fri <br> 10.25 | TEST- cumulative | TEST- cumulative |

## WARM UP ACTIVITIES

| $\begin{aligned} & \text { MON } \\ & 10.21 \end{aligned}$ | Describe 2-3 genetic traits that you are SURE came from just 1 of your parents. <br> Write balanced equations for the following: 1-Hydrogen gas burns in air in a synthesis reaction <br> 2-Zinc metal is placed in sulfuric acid (single displacement) <br> 3-Ammonium chloride reacts with Magnesium Phosphate in a double displacement reaction. |
| :---: | :---: |
| TUE $10.22$ | Which numbers did you get wrong? (list here). Do corrections on the back of this sheet for homework © BDACC, BEBBA <br> Write and balance the following equations using Calcium metal rather than sodium metal: |


|  | 1) $\mathrm{Na}+\mathrm{Cl}_{2} \rightarrow 2 \mathrm{NaCl}$ <br> 2) $2 \mathrm{Na}+2 \mathrm{H}_{2} \mathrm{O} \rightarrow 2 \mathrm{NaOH}+\mathrm{H}_{2}$ <br> 3) $2 \mathrm{Na}+2 \mathrm{HCl} \rightarrow 2 \mathrm{NaCl}+\mathrm{H}_{2}$ <br> 4) $\mathrm{Na}_{2} \mathrm{CO}_{3} \rightarrow \mathrm{Na}_{\mathbf{2}} \mathrm{O}+\mathbf{C O}_{2}$ <br> 5) $2 \mathrm{NaNO}_{3}+\mathbf{H}_{2} \mathrm{SO}_{4} \rightarrow \mathrm{Na}_{2} \mathrm{SO}_{4}+2 \mathrm{HNO}_{3}$ |
| :---: | :---: |
| $\begin{aligned} & \text { WED } \\ & 10.23 \end{aligned}$ | A plant that is heterogygous tall is pollinated by a plant that is homozygous tall. TALL IS DOMINANT. What are the chances of getting a short F1 generation? <br> Which numbers did you get wrong? (list here): Do corrections on the back of this sheet for homework ${ }^{()}$ |
| $\begin{aligned} & \text { THU } \\ & 10.24 \end{aligned}$ | 1)HAND IN ENTIRE PACKET!! 2) Describe the difference between Incomplete dominance and co-dominance. <br> Which numbers did you get wrong? (list here): Do corrections on the back of this sheet for homework © |
| $\begin{aligned} & \text { FRI } \\ & 10.25 \end{aligned}$ | Give three examples of sex-linked disorders....then HAND IN WARM UPS!! Is breathing a chemical change? Explain why or why not....then HAND IN WARM UPS!! |

## The Law of Dominance

## Stated "simply" it goes like so:

In a cross of parents that are pure for contrasting traits, only one form of the
trait will appear in the next generation. Offspring that are hybrid for a trait will have only the dominant trait in the phenotype.

## The Law of Segregation

Goes like so: During the formation of gametes (eggs or sperm), the two alleles responsible for a trait separate from each other. Alleles for a trait are then "recombined" at fertilization, producing the genotype for the traits of the offspring.

## The Law of Independent Assortment

Alleles for different traits are distributed to sex cells (\& offspring) independently of one another.



https://hobart.k12.in.us/jkousen/Biology/mendel.htm

1. O When several genes influence a trait.
2. $\vee$ a version of a gene
3. $N$ When the heterozygous genotype results in a phenotype where both alleles are fully and separately expressed
4. A Chromosomes 1-44 in a human
5. E Only expressed in the homozygous state
6. Q Genes that travel on the X chromosome
7. $\llcorner$ Chromosomes line up randomly during metaphase therefore it is possible for any combination of chromosomes to be passed on from parent to offspring.
8. $T$ A diagram that shows homologous chromosome pairs
9. $W$ The macromolecule that runs your body and expresses your traits
10. G A genotype resulting from the inheritance of two different alleles from your parents
11. $Y$ Section of a chromosome that codes for a single protein
12. $R$ Genes that are likely inherited together due to their physical proximity
13. D Overshadows the other allele in the heterozygous state
14. H The physical trait expressed
15. P More than 2 versions of a gene (more than just a "dominant" and a "recessive")
16. $C$ Condensed genetic material
17. $S$ When one gene overshadows another
18. K At the end of meiosis, each gamete formed should only have 1 copy form each homologous chromosome pair
19. F A genotype resulting from the inheritance of the same alleles from your parents
20. X The macromolecule that has the instructions for making you who you are
21. $M$ when the heterozygous genotype results in a phenotype where the two alleles are blended together
22. I The actual alleles you inherit
23. $B$ The chromosomes that determine your sex
24. $\checkmark$ A person that has the gene for a trait or disease but doesn't show it
25. J Some versions of genes are dominant over others.

## Practice: Complex Inheritance Word Problems

Answer the questions below about different complex inheritance patterns. Be sure to include a Punnett square to support your answers for \#1-5.

1. In chickens, black feathers are codominant to white feathers. Create a Punnett square for a cross between two chickens that have both black and white feathers. Then list the phenotypic and genotypic
ratios of offspring.
$25 \%$


$$
\begin{array}{ll}
\text { Geno }=1: 2: 1 & 50 \% \\
& 25 \% \\
\text { Phone }=1: 2: 1 & 25 \% \text { Black } \\
& 50 \% \text { Specked } \\
& 25 \% \text { White }
\end{array}
$$

2. Tim and Stephanie are devastated when they find out their newborn son has hemophilia - a sex-linked recessive disease. Tim is shocked becaused he doesn't have hemophilia, and figures if his son has it he would have gotten it from him. Is Tim correct in his thinking? Explain. Also explain what their genotypes must be in order for neither of them to have the disease but have a son who does.


- no b/c their son got Y from dad + affected $X$ from mom.
- Tim $=X^{H} Y$

Steph : $X^{H} X^{h}$
3. Tim and Stephanie (from \#2) are pregnant again! This time with a baby girl. They are nervous that she too may get hemophilia. Based on what you know about Tim and Stephanie, what is the probability that they do have a daughter with the disease? Show a Punnett square to support your answer.

4. In carnations, red and white flowers make pink flowers. What complex inheritance pattern is this? Also, create a Punnett square for a cross between a red flower and a pink flower. Then list the phenotypic and genotypic rations of the offspring.

$$
\begin{aligned}
& \text { incomplete dominance }
\end{aligned}
$$

5. Jessica is blood type A and her husband Graham is blood type O. She is worried because they are pregnant and their future child has a rare disorder which will require him to need regular blood transfusions. Jessica worries that if their son ends up with type O blood, like his father, it will be hard to get the necessary regular blood transfusions because people with O blood can only accept O blood. What is the likelihood of their child having blood type O? Use Punnett squares to show all possibilities.

$0 \%$ if Jessica is homozygous type A
$S 0 \%$ if Jessica is hetrozygous Type A
6. In Labrador retrievers, some puppies have pink noses and some have black. Labrador retrievers with black fur almost always have black noses. What type of inheritance pattern is this? Explain how this is possible.

> Linked genes
> possible if gene for tor color + nose color are physically close on the same chromosome.
7. In cats, some are black, some are orange, and some are calico (fur that is both black and orange.) All calico cats are always female. This trait represents two inheritance patterns. List which two and explain how you know.

$$
\begin{aligned}
& \text { codominance b/c black }+ \text { orange } \\
& \text { sex-linked b/c only females get it } X^{+1} X^{n}
\end{aligned}
$$

8. An organisms has three different versions of Gene Pyx - version $\mathrm{a}, \mathrm{b}$ and c . Determine the inheritance pattern of this trait. Then list below all of the different genotypes that an organism with this gene could have.

| Multiple alleles |  |
| :---: | :---: |
| $a a$ | $a b$ |
| $b b$ | $a c$ |
| $c c$ | $b c$ |



## 1. SYNTHESIS:

a. Formation of binary compound: $A+B \rightarrow A B$
b. Metal oxide-water reactions: $\mathrm{MO}+\mathrm{H} 2 \mathrm{O} \rightarrow$ base
c. Nonmetal oxide-water reactions: (NM)O $+\mathrm{H} 2 \mathrm{O} \rightarrow$ acid
2. DECOMPOSITION:
a. Binary compounds: $\mathrm{AB} \rightarrow \mathrm{A}+\mathrm{B}$
b. Metallic carbonates: $\mathrm{MCO} \rightarrow \mathrm{MO}+\mathrm{CO} 2$
c. Metallic hydrogen carbonates: $\mathrm{MHCO} \rightarrow \mathrm{MO}+\mathrm{H} 2 \mathrm{O}(\mathrm{I})+\mathrm{CO} 2(\mathrm{~g})$
d. Metallic hydroxides: $\mathrm{MOH} \rightarrow \mathrm{MO}+\mathrm{H} 2 \mathrm{O}$
e. Metallic chlorates: $\mathrm{MClO} 3 \rightarrow \mathrm{MCl}+\mathrm{O} 2$
f. Oxyacids decompose to nonmetal oxides and water: acid $\rightarrow(\mathrm{NM}) \mathrm{O}+\mathrm{H} 2 \mathrm{O}$
3. SINGLE REPLACEMENT:
a. Metal-metal replacement: $A+B C \rightarrow A C+B$
b. Active metal replaces H from water: $\mathrm{M}+\mathrm{H} 2 \mathrm{O} \rightarrow \mathrm{MOH}+\mathrm{H} 2$
c. Active metal replaces H from acid: $\mathrm{M}+\mathrm{HX} \rightarrow \mathrm{MX}+\mathrm{H} 2$
d. Halide-Halide replacement: $D+B C \rightarrow B D+C$
4. DOUBLE REPLACEMENT: $A B+C D \rightarrow A D+C B$
a. Formation of a precipitate from solution
b. Acid-Base neutralization reaction
5. COMBUSTION REACTION Hydrocarbon + oxygen $\rightarrow$ carbon dioxide + water

Synthesis
a) $2 \mathrm{Al}+3 \mathrm{~S} \rightarrow \mathrm{Al}_{2} \mathrm{~S}_{3}$
b) $\mathrm{Na}_{2} \mathrm{O}+\mathrm{H}_{2} \mathrm{O} \rightarrow 2 \mathrm{NaOH}$
c) $\mathrm{SO}_{3}+\mathrm{H}_{2} \mathrm{O} \rightarrow \mathrm{H}_{2} \mathrm{SO}_{4}$

Decomposition
a) $2 \mathrm{H}_{2} \mathrm{O} \rightarrow 2 \mathrm{H}_{2}+\mathrm{O}_{2}$
b) $\mathrm{Li}_{2} \mathrm{CO}_{3} \rightarrow \mathrm{Li}_{2} \mathrm{O}^{2}+\mathrm{CO}_{2}$
c) $\mathrm{NaHCO}_{3} \rightarrow \mathrm{Na}_{2} \mathrm{O}+\mathrm{H}_{2} \mathrm{O}+\mathrm{CO}_{2}$
d) $2 \mathrm{NaOH} \rightarrow \mathrm{Na}_{2} \mathrm{O}+\mathrm{H}_{2} \mathrm{O}$
e) $2 \mathrm{NaClO}_{3} \rightarrow 2 \mathrm{NaCl}+3 \mathrm{O}_{2}$
f)

Single replacement
a) $\mathrm{Na}+\mathrm{LiCl} \rightarrow \mathrm{NaCl}+\mathrm{Li}$
b) $2 \mathrm{Na}+\mathrm{H}_{2} \mathrm{O} \rightarrow 2 \mathrm{NaOH}+\mathrm{H}_{2}$
c) $2 \mathrm{Na}+2 \mathrm{HCl} \rightarrow 2 \mathrm{NaCl}+\mathrm{H}_{2}$
d) $\mathrm{F}_{2}+2 \mathrm{HBr} \rightarrow \mathrm{B}_{2}+2 \mathrm{HF}$

Double replacement
a) $\mathrm{Pb}\left(\mathrm{NO}_{3}\right)_{2}+2 \mathrm{KI} \rightarrow \mathrm{PbI}_{2(\mathrm{~s})}+2 \mathrm{KNO}_{3}$
b)

Combustion
a) $\mathrm{C}_{2} \mathrm{H}_{6}+\mathrm{O}_{2} \rightarrow$
b) $\mathrm{C}_{2} \mathrm{H}_{2}+\mathrm{O}_{2} \rightarrow$
c) $\mathrm{C}_{3} \mathrm{H}_{8}+\mathrm{O}_{2} \rightarrow$
d) $\mathrm{C}_{4} \mathrm{H}_{10}+\mathrm{O}_{2} \rightarrow$


## A GUIDE TO THE TWENTY COMMON AMINO ACIDS

AMINO ACIDS ARE THE BUILDING BLOCKS OF PROTEINS IN LIVING ORGANISMS. THERE ARE OVER 500 AMINO ACIDS FOUND IN NATURE - HOWEVER, THE HUMAN GENETIC CODE ONLY DIRECTLY ENCODES 20. 'ESSENTIAL' AMINO ACIDS MUST BE OBTAINED FROM THE DIET, WHILST NON-ESSENTIAL AMINO ACIDS CAN BE SYNTHESISED IN THE BODY.
Chart Key: aliphatic aromatic acidic basic hydroxylic sulfur-containing amidic $\bigcirc$ non-essential indential


TYROSINE (Y)
$\underset{\substack{\text { Tyr } \\ \text { tart tac }}}{ }$




GLUTAMICACID E


Met
atG


Pro
cct, coc.cca, ccg


Asn

$\operatorname{cin}_{6}^{6}$

Note: This chart only shows those amino acids for which the human genetic code directly codes for. Selenocysteine is often referred to as the 21 st amino acid, but is encoded in a special manner In some cases, distinguishing between asparagine/aspartic acid and glutamine/glutamic acid is difficult. In these cases, the codes asx (B) and glx (Z) are respectively used.

## ALWAYS show all work and units and units of WHAT

| -Iron III means $\mathrm{Fe}^{3+}$ oxygen is diatomic -Be sure to match up charges correctly to make compounds! -hydrogen gas is diatomic | -copper II means Cu² <br> -the pentagon has 5 <br> sides <br> -find \% water in the whole thing <br> Mass water/mass whole <br> thing $\mathrm{x} 100=\%$ water |
| :---: | :---: |
| assume out of 100 g and convert all to moles! -find the mole ratio | - Use units and units of WHAT for everything. \% yield= <br> ACTUAL /THEORETICAL x 100 |

## Part 5- $\quad \mathbf{N H}_{3}$ is the product



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